

UNDERSTANDING DEFECTS AND
EXCITED-STATE PROPERTIES IN
PEROVSKITES FROM MANY-BODY
PERTURBATION THEORY

MANISH KUMAR



DEPARTMENT OF PHYSICS
INDIAN INSTITUTE OF TECHNOLOGY DELHI
JULY 2022

© Indian Institute of Technology Delhi (IITD), New Delhi, 2022

UNDERSTANDING DEFECTS AND EXCITED-STATE PROPERTIES IN PEROVSKITES FROM MANY-BODY PERTURBATION THEORY

by

MANISH KUMAR

Department of Physics

Submitted

in fulfillment of the requirements of the degree of Doctor of Philosophy

to the



INDIAN INSTITUTE OF TECHNOLOGY DELHI

JULY 2022

Dedicated to my parents

Certificate

This is to certify that the thesis entitled “**Understanding Defects and Excited-State Properties in Perovskites from Many-Body Perturbation Theory**” being submitted by **Manish Kumar**, to the Indian Institute of Technology Delhi, for the award of the degree of **Doctor of Philosophy** in Physics is a record of bonafide research work carried out by him under my supervision and guidance. He has fulfilled the requirements for the submission of the thesis, which to the best of my knowledge has reached the required standard. The material contained in the thesis has not been submitted in part or full to any other University or Institute for the award of any degree or diploma.

Prof. Saswata Bhattacharya

Thesis Supervisor

Department of Physics,

Indian Institute of Technology Delhi,

Hauz Khas, New Delhi 110016, India.

Date:.....

Place: New Delhi

Acknowledgments

I would like to express my deepest gratitude and sincere appreciation to my thesis supervisor Prof. Saswata Bhattacharya for his inspiring and ever-encouraging guidance. His scientific temperament, logical thinking, scientist intuition, passion for physics, expertise and enthusiasm has been of great value to me and will have a bearing for the rest of my life. His invaluable advises, comments, new ideas and research guidance has been the cornerstone of my growth as a student and an independent research scholar. I would like to recognize his invaluable assistance in improving my presentation and scientific writing skills. I am extremely grateful to him for his relentless support during my PhD journey from day one. His immense patience and willingness to help everyone in any regard have had a profound effect on me.

I would like to thank all my colleagues and collaborators from our vibrant research group DISCERE (DISCovering Electronic CorRELation): Dr. Shikha Saini, Dr. Pooja Basera, Dr. Ekta Arora, Arunima Singh, Deepika Gill, Manjari Jain, Preeti Bhumla, Sajjan Sheoran and Ankita Phutela for their constant help and support. Their constant motivation, love and affection helped me to improve a lot and reach this stage. A very special thanks to Dr. Pooja Basera for several fruitful discussions and assistance throughout the whole period of my research work.

I would also like to thank my other collaborators: Prof. Venkat Krishnan, Prof. G. Vijay Prakash, Prof. Ritu Gupta, Prof. M. Ali Haider, Dr. Ashish Kumar, Dr. Sonit Balyan and Kshetra Mohan Dehury for useful discussions and collaborations.

I am deeply grateful to the Indian Institute of Technology Delhi (IIT Delhi) for providing me all the facilities to carry out my research work and providing the travel grant for participating in conferences. I am thankful to Council for Scientific and Industrial Research (CSIR) for funding my research and travel for conferences.

I would like to take this opportunity to thank my student research committee members: Prof. Sankalpa Ghosh, Prof. B. K. Mani and Prof. Hemant K. Kashyap for their evaluation of my

research work from time to time. Their intriguing questions and invaluable comments have guided me to look deeper into the problems. I would like to extend my sincere thanks to Prof. Amrita Bhattacharya for insightful discussions.

A word of sincere appreciation and deep gratitude for all my teachers for their blessings, encouragement and motivation that helped me to get this far. I sincerely thank all my friends here at IIT Delhi, especially Sandeep, Arvind, Sajjan, Hemant, Chandan, Aditya and Sooryansh for making my tenure at the campus all the more memorable and outside the campus, especially Nishant, Vishnu and Prerna for their consistent encouragement. I apologize for my failure to mention everyone by name.

Most importantly, I would like to thank my family. This thesis would not be possible without the unconditional love, support and encouragement of my parents. I appreciate all of the struggles and sacrifices that they have made for me. I am so blessed to have them in my life.

Manish Kumar

Abstract

Oxide perovskites such as SrTiO_3 and CaTiO_3 are the potential candidates to be used as a photocatalyst due to their exceptional electronic structure, high chemical stability, non-toxicity, and low cost. They exhibit suitable conduction and valence band-edge positions for reduction and oxidation of water to produce hydrogen and oxygen. Therefore, they can be exploited to generate hydrogen via water splitting, which is a clean, sustainable, and abundant source of energy. However, owing to their wide band gap, they absorb only UV irradiation (which consists $\sim 4\%$ of the solar spectrum). Hence, several works are dedicated to expand optical response toward the visible region by reducing the band gap through doping with metals, nonmetals, or the combination of different elements. Despite significant amount of research is done, both experimentally and theoretically on these systems, it is still an open question concerning the kind of dopants or codopants, that could reduce the band gap while retaining the photocatalytic efficiency. In view of this, we systematically study the role of monodoping as well as codoping of a metal and nonmetal in SrTiO_3 in enhancing the photocatalytic efficiency for water splitting. Moreover, we investigate the effects of intrinsic defect (viz. O-vacancy) in CaTiO_3 to disentangle the role of O-vacancy for water splitting and N_2 fixation reaction.

On the other hand, lead halide perovskites have emerged as an efficient compound semiconductor alternative to conventional materials used in photovoltaics. This class of materials has suitable optical band gap, long carrier diffusion length, high charge carrier mobility and low manufacturing cost. However, the concerns regarding toxicity of lead and phase instability restricts their usage on large scale. In an attempt to deal with toxicity and instability, lead-free halide double perovskites such as $\text{Cs}_2\text{M(I)M(III)X}_6$ ($\text{M} = \text{metal}$, $\text{X} = \text{halogen}$) and chalcogenide perovskites ABX_3 ($\text{A}, \text{B} = \text{metals}$, $\text{X} = \text{chalcogen}$) have emerged. In this work, we aim to design lead-free halide double perovskites with improved optoelectronic properties since they have not shown the efficiency as that of lead halide perovskites. Furthermore, the excitonic and polaronic effects are unraveled in the case of chalcogenide perovskites.

We employ a robust methodological approach that integrates various levels of theories combined into one multi-scale simulation to address the optical properties such as dielectric function, absorption spectra, exciton binding energy and polaronic effects in perovskites. In this thesis work, the state-of-the-art methodologies that are used to obtain the desired objectives are: (i) density functional theory (DFT) for ground-state properties, (ii) *ab initio* atomistic thermodynamics to predict the stability, (iii) many-body perturbation theory (GW, BSE and model-BSE) for excited-state properties, (iv) Wannier-Mott approach to determine the exciton binding energy and exciton lifetime, and (v) density functional perturbation theory (DFPT) for including ionic contribution to dielectric function and to capture the electron-phonon coupling.

ऑक्साइड पैरोस्काइट्स जैसे कि SrTiO_3 और CaTiO_3 सामर्थ्यवान उम्मीदवार हैं जिन्हें उनकी असाधारण इलेक्ट्रॉनिक संरचना, उच्च रासायनिक स्थिरता, गैर-विषाक्तता और कम लागत के कारण फोटोकैटलिस्ट के रूप में उपयोग किया जा सकता है। वे हाइड्रोजन और ऑक्सीजन का उत्पादन करने के लिए पानी का अपचयन और ऑक्सीकरण के लिए उपयुक्त चालन और संयोजकता बैंड-एज स्थिति प्रदर्शित करते हैं। इसलिए, पानी के विभाजन के माध्यम से हाइड्रोजन उत्पन्न करने के लिए उनका दोहन किया जा सकता है, जो ऊर्जा का एक स्वच्छ, टिकाऊ और प्रचुर स्रोत है। हालांकि, उनके व्यापक बैंड अंतराल के कारण, वे केवल यूवी विकिरण (जो सौर स्पेक्ट्रम का लगभग 4% होता है) को अवशोषित करते हैं। इसलिए, धातुओं, अधातुओं या विभिन्न तत्वों के संयोजन के साथ मादन (डोपिंग) के माध्यम से बैंड अंतराल को कम करके दृश्य क्षेत्र की ओर ऑप्टिकल प्रतिक्रिया का विस्तार करने के लिए कई कार्य समर्पित हैं। इन प्रणालियों पर प्रायोगिक और सैद्धांतिक दोनों तरह से विस्तृत मात्रा में शोध किए जाने के बावजूद, यह अभी भी एक खुला प्रश्न है कि किस तरह के अपमिश्रक (डोपेंट) या कोडोपेंट हैं, जो फोटोकैटलिटिक दक्षता को बनाए रखते हुए बैंड अंतराल को कम कर सकते हैं। इसे ध्यान में रखते हुए, हम पानी के विभाजन के लिए फोटोकैटलिटिक दक्षता को बढ़ाने में SrTiO_3 में मोनोडोपिंग के साथ-साथ धातु और अधातु के कोडोपिंग की भूमिका का व्यवस्थित रूप से अध्ययन करते हैं। इसके अलावा, हम पानी के विभाजन और N_2 निर्धारण प्रतिक्रिया के लिए O-रिक्ति की भूमिका को सुलझाने के लिए CaTiO_3 में आंतरिक दोष (अर्थात् O-रिक्ति) के प्रभावों की जांच करते हैं।

दूसरी ओर, लैंड हैलाइड पैरोस्काइट्स एक कुशल यौगिक अर्धचालक फोटोवोल्टिक में प्रयुक्त पारंपरिक सामग्रियों के विकल्प के रूप में उभरे हैं। सामग्री के इस वर्ग में उपयुक्त ऑप्टिकल बैंड अंतराल, लंबी वाहक प्रसार लंबाई, उच्च आवेश वाहक गतिशीलता और कम विनिर्माण लागत है। हालांकि, लैंड की विषाक्तता और चरण अस्थिरता के बारे में चिंताएं बड़े पैमाने पर उनके उपयोग को प्रतिबंधित करती हैं। विषाक्तता और अस्थिरता से निपटने के प्रयास में, लैंड-रहित हैलाइड डबल पैरोस्काइट्स जैसे $\text{Cs}_2\text{M(I)M(III)X}_6$ (M = धातु, X = हैलोजन) और चाल्कोजेनाइड पैरोस्काइट्स ABX_3 (A, B = धातु, X = चाल्कोजेन) उभरे हैं। इस काम में, हमारा लक्ष्य बेहतर ऑप्टोइलेक्ट्रॉनिक गुणों के साथ लैंड-रहित हैलाइड डबल पैरोस्काइट्स को डिज़ाइन करना है क्योंकि उन्होंने लैंड हैलाइड पैरोस्काइट्स जैसी दक्षता नहीं दिखाई है। इसके अलावा, चाल्कोजेनाइड पैरोस्काइट्स के मामले में एक्साइटोनिक और ध्रुवीय प्रभाव सुलझाये हैं।

हम एक मजबूत कार्यप्रणाली दृष्टिकोण को नियोजित करते हैं जो ऑप्टिकल गुणों जैसे कि परावैद्युत फलन, अवशोषण स्पेक्ट्रा, एक्साइटन बाध्यकारी ऊर्जा और पैरोस्काइट्स में ध्रुवीय प्रभाव को संबोधित करने के लिए एक बहु-स्तरीय अनुरूपण में संयुक्त सिद्धांतों के विभिन्न स्तरों को एकीकृत करता है। इस शोध-प्रबन्ध कार्य में, वांछित उद्देश्यों को प्राप्त करने के लिए उपयोग की जाने वाली अत्याधुनिक पद्धतियां हैं: (i) ग्राउंड-स्टेट गुणों के लिए घनत्व कार्यात्मक सिद्धांत (डीएफटी), (ii) स्थिरता की भविष्यवाणी करने के लिए आदित परमाणु ऊष्मा गतिकी, (iii) एक्साइटड-स्टेट गुणों के लिए बहुपिंडी क्षोभ सिद्धांत (जीडब्ल्यू, बीएसई और मॉडल-बीएसई), (iv) एक्साइटॉन बाध्यकारी ऊर्जा और एक्साइटॉन जीवनकाल निर्धारित करने के लिए वैनियर-मॉट दृष्टिकोण, और (v) घनत्व कार्यात्मक क्षोभ सिद्धांत (डीएफपीटी) परावैद्युत फलन में आयनिक योगदान को शामिल करने और इलेक्ट्रॉन-फोनॉन युग्मन के अध्ययन के लिए।

List of Publications

1. **Manish Kumar**, Pooja Basera, Shikha Saini, and Saswata Bhattacharya, "Role of defects in photocatalytic water splitting: Monodoped vs codoped SrTiO₃", *The Journal of Physical Chemistry C* **124**, 10272 (2020).
2. **Manish Kumar**, Pooja Basera, Shikha Saini, and Saswata Bhattacharya, "Theoretical insights of codoping to modulate electronic structure of TiO₂ and SrTiO₃ for enhanced photocatalytic efficiency", *Scientific Reports* **10**, 15372 (2020).
3. **Manish Kumar**, Manjari Jain, Arunima Singh, and Saswata Bhattacharya, "Sublattice mixing in Cs₂AgInCl₆ for enhanced optical properties from first-principles", *Applied Physics Letters* **118**, 021901 (2021).
4. **Manish Kumar**, Arunima Singh, Deepika Gill and Saswata Bhattacharya, "Optoelectronic properties of chalcogenide perovskites by many-body perturbation theory", *The Journal of Physical Chemistry Letters* **12**, 5301 (2021).
5. Ashish Kumar, **Manish Kumar**, Navakoteswara Rao, Muthukonda Venkatakrishnan Shankar, Saswata Bhattacharya, and Venkata Krishnan, "Unraveling the structural and morphological stability of oxygen vacancy engineered leaf-templated CaTiO₃ towards photocatalytic H₂ evolution and N₂ fixation reactions", *Journal of Materials Chemistry A* **9**, 17006 (2021).
6. Ekta Arora, Shikha Saini, Pooja Basera, **Manish Kumar**, Arunima Singh, and Saswata Bhattacharya, "Elucidating the role of temperature and pressure to the thermodynamic stability of charged defects in complex metal-hydrides: A case study of NaAlH₄", *The Journal of Physical Chemistry C* **123**, 62 (2019).
7. Pooja Basera, Shikha Saini, Ekta Arora, Arunima Singh, **Manish Kumar**, and Saswata Bhattacharya, "Stability of non-metal dopants to tune the photo-absorption of TiO₂ at

- realistic temperatures and oxygen partial pressures: A hybrid DFT study”, Scientific Reports **9**, 11427 (2019).
8. Arunima Singh, Pooja Basera, Shikha Saini, **Manish Kumar**, and Saswata Bhattacharya, “Importance of many-body dispersion in the stability of vacancies and antisites in free-standing monolayer of MoS₂ from first-principles approaches”, The Journal of Physical Chemistry C **124**, 1390 (2020).
 9. Pooja Basera, **Manish Kumar**, Shikha Saini, and Saswata Bhattacharya, “Reducing lead toxicity in the methylammonium lead halide MAPbI₃: Why Sn substitution should be preferred to Pb vacancy for optimum solar cell efficiency”, Physical Review B **101**, 054108 (2020).
 10. Manjari Jain, Arunima Singh, Pooja Basera, **Manish Kumar**, and Saswata Bhattacharya, “Understanding the role of Sn-substitution and Pb-□ in enhancing the stability of CH(NH₂)₂Pb_{1-X-Y}Sn_X□_YBr₃: A hybrid density functional approach”, Journal of Materials Chemistry C **8**, 10362 (2020).
 11. Deepika Gill, **Manish Kumar**, Pooja Basera, and Saswata Bhattacharya, “Understanding the ionic diffusivity in (meta)stable (un)doped solid state electrolyte from first-principles: A case study of LISICON”, The Journal of Physical Chemistry C **124**, 17485 (2020).
 12. Gaurav Bahuguna, Indrajit Mondal, Mohit Verma, **Manish Kumar**, Saswata Bhattacharya, Ritu Gupta, and Giridhar U. Kulkarni, “Innovative approach to photo-chemiresistive sensing technology: Surface-fluorinated SnO₂ for VOC detection”, ACS Applied Materials & Interfaces **12**, 37320 (2020).
 13. Shikha Saini, Pooja Basera, **Manish Kumar**, Preeti Bhumla, and Saswata Bhattacharya, “Metastability triggered reactivity in clusters at realistic conditions: A case study of N-doped (TiO₂)_n for photocatalysis”, Journal of Physics Materials **4**, 015001 (2020).
 14. Preeti Bhumla, **Manish Kumar**, and Saswata Bhattacharya, “Theoretical insights into C–H bond activation of methane by transition metal clusters: The role of anharmonic effects”, Nanoscale Advances **3**, 575 (2021).
 15. Deepika Gill, Preeti Bhumla, **Manish Kumar**, and Saswata Bhattacharya, “High-throughput

-
- screening to modulate electronic and optical properties of alloyed $\text{Cs}_2\text{AgBiCl}_6$ for enhanced solar cell efficiency”, *Journal of Physics Materials* **4**, 025005 (2021).
16. Kshetra Mohan Dehury, Pawan K. Kanaujia, Mohammad Adnan, **Manish Kumar**, Saswata Bhattacharya, and G. Vijaya Prakash, “Structure-dependent (non)linear optical excitons in primary cyclic ammonium ($\text{C}_n\text{H}_{2n-1}\text{NH}_2$; $n = 3 - 8$)-based inorganic-organic hybrid semiconductor series”, *The Journal of Physical Chemistry C* **125**, 6821 (2021).
17. Sajjan Sheoran, **Manish Kumar**, Preeti Bhumla, and Saswata Bhattacharya, “Rashba spin splitting and anomalous spin textures in the bulk ferroelectric oxide perovskite KIO_3 ”, *Materials Advances* **3**, 4170 (2022).
18. Manjari Jain, Preeti Bhumla, **Manish Kumar** and Saswata Bhattacharya, “Lead-free alloyed double perovskites: An emerging class of materials for optoelectronic applications”, *The Journal of Physical Chemistry C* **126**, 6753 (2022).

Contents

Certificate	i
Acknowledgements	ii
Abstract	iv
List of Publications	viii
List of Figures	xv
List of Tables	xxi
1 Introduction	1
1.1 Defects in solids	1
1.2 Thermodynamics of point defects	2
1.3 Defect dependent properties	3
1.4 Defects in perovskites	5
1.5 Problems and challenges	8
1.6 A short overview of this thesis	10
2 Theoretical methodology	13
2.1 Computer simulation	13
2.2 First-principles calculation	15
2.3 Many-body physics: A theoretical framework	15
2.4 Time-independent many-body Schrödinger equation	17
2.4.1 The Hartree approximation	19
2.4.2 The Hartree–Fock approximation	20
2.5 Density functional theory (DFT)	21

2.5.1	Thomas-Fermi-Dirac approximation	22
2.5.2	The Hohenberg-Kohn theorems	23
2.5.3	The Kohn-Sham ansatz	25
2.5.4	Exchange-correlation functionals	26
2.5.4.1	Local Density Approximation (LDA)	26
2.5.4.2	Generalized Gradient Approximation (GGA)	28
2.5.4.3	Meta-Generalized Gradient Approximation (meta-GGA)	30
2.5.4.4	Hybrid functionals	30
2.6	Basis set	31
2.6.1	Plane waves basis set	32
2.6.2	Pseudopotentials	36
2.6.3	Norm-conserving pseudopotentials	38
2.6.4	Ultrasoft pseudopotentials	39
2.6.5	Projector augmented-wave (PAW) method	40
2.7	Force theorem and geometry optimization	41
2.8	<i>Ab initio</i> atomistic thermodynamics	42
2.8.1	Thermodynamic potentials	42
2.8.2	Defect formation energy	43
2.8.3	Chemical potentials	43
2.9	Many-body perturbation theory (MBPT): The Green's function approach	46
2.9.1	Green's function	47
2.9.2	Dyson's equation: The self-energy operator Σ	50
2.9.3	Hedin's equations and the <i>GW</i> approximation	51
2.9.4	Practical implementation of the single-shot <i>GW</i> (G_0W_0)	52
2.9.5	Bethe-Salpeter equation (BSE)	53
2.9.6	Optical Spectrum	57
2.10	Density functional perturbation theory (DFPT)	58
2.10.1	Lattice dynamics from electronic structure theory	58
2.10.2	Linear response	59
3	Role of defects in photocatalytic water splitting: Monodoped vs codoped SrTiO₃	61
3.1	Introduction	61
3.2	Computational methods	64

3.3	Results and discussion	66
3.3.1	Stability of defects in SrTiO ₃ : <i>Ab initio</i> atomistic thermodynamics	66
3.3.1.1	N-related defects	69
3.3.1.2	Mn-related defects	71
3.3.1.3	Codoped SrTiO ₃	71
3.3.2	Electronic structure analysis	75
3.3.3	Optical properties	77
3.3.4	Band-edge alignment	80
3.3.5	Band structure and effective mass of pristine, Mn _{Sr} N _O , and Mn _{Ti} S _O codoped SrTiO ₃	82
3.4	Conclusions	84
4	Unraveling the role of oxygen vacancy in CaTiO₃ for photocatalytic applications	85
4.1	Introduction	85
4.2	Computational methods	87
4.3	Results and discussion	88
4.3.1	Electronic structure of (un)defective CaTiO ₃	88
4.3.2	H ₂ evolution in (un)defective CaTiO ₃ from photocatalytic water splitting	89
4.3.3	N ₂ fixation in (un)defective CaTiO ₃	90
4.4	Conclusions	91
5	Sublattice mixing in Cs₂AgInCl₆ for enhanced optical properties from first-principles	93
5.1	Introduction	93
5.2	Computational methods	94
5.3	Results and discussion	96
5.3.1	Stability of defected systems	96
5.3.1.1	Structural stability	96
5.3.1.2	Thermodynamic stability	98
5.3.2	Electronic structure analysis	104
5.3.3	Optical properties	108
5.4	Conclusions	112

6	Optoelectronic properties of chalcogenide perovskites by many-body perturbation theory	114
6.1	Introduction	114
6.2	Computational methods	115
6.3	Results and discussion	118
6.3.1	Electronic structure	118
6.3.2	Optical properties	121
6.3.3	Polaronic effects	126
6.3.4	Theoretical efficiency	128
6.4	Conclusions	129
7	Epilogue and outlook	131

List of Figures

1.1	Schematic illustration of (a) point (e.g., vacancy, substitutional, and interstitial), (b) line (e.g., edge dislocation), and (c) surface (e.g., defect at grain boundaries) defects.	1
1.2	Schematic illustration of shallow and deep defect states. Here, E_g is the band gap of the pristine material.	4
1.3	Schematic illustration of perovskites and their applications.	5
1.4	Schematic representation of formation of halide double perovskites $A_2B(I)B'(III)X_6$ to exclude Pb from LHPs $APbX_3$. The A^+ , Pb(II), B(I), B'(III), X^- ions are denoted by dark red, light blue, light green, orchid, and golden color balls, respectively.	7
1.5	Schematic illustration of the proposed strategies to design and study thermodynamic stability, electronic and optical properties of semiconducting perovskites for various applications.	8
2.1	Multi-scale simulation in various length and time scales.	14
2.2	Schematic representation of mapping of interacting system to a non-interacting many-electron system through the same ground-state electron density.	25
2.3	Flow chart to solve the Kohn-Sham equations self-consistently.	27
2.4	Jacob's ladder of density functional approximations [1]	29
2.5	Schematic representation of pseudopotential technique. The all-electron wave function corresponding to Coulomb potential is shown by red color. The pseudo wave function corresponding to pseudopotential is shown by blue color.	37

- 2.6 Schematic representation of defect formation energy as a function of chemical potential of electron at a particular (T, p) , which can exist in three charge states $q = 0, +1$, and -1 . $\varepsilon(+/0)$ and $\varepsilon(0/-)$ are the charge-state transition levels, denoting a deep donor level and a deep acceptor level, respectively. The thick green colored lines indicate the most favorable charge state for a given value of μ_e 44
- 2.7 Schematic representation of excited-state spectroscopies, namely, direct photoemission, inverse photoemission, and optical absorption. Here, IP and EA represent the ionization potential and electron affinity, respectively. Also, E_N is the total energy of N -electron system. Moreover, $E_g^{\text{GW}} = \text{IP} - \text{EA}$ is the quasiparticle (QP) band gap and $E_g^{\text{BSE}} = \text{IP} - \text{EA} - E_B$ is the optical band gap, where E_B is the exciton binding energy. 47
- 2.8 Schematic representation of spectral function in the case of non-interacting (electrons) single-particle excitation and interacting single-particle like (QP) excitation. 48
- 2.9 Illustration of a QP formation in the case of photoemission spectroscopy. 49
- 2.10 Schematic representation of the Dyson's equation, which relates the non-interacting (G_0) and interacting (G) Green's functions via the self-energy operator (Σ). Here, the black arrow describes the propagation of a non-interacting particle and the red color represents screening process of different orders. 50
- 2.11 Schematic representation of the self-consistent Hedin's equations. 52
- 3.1 The formation energy of a single O-vacancy defect as a function of chemical potential of electron under O-rich condition using (a) LDA, (b) PBE, and (d) HSE06 ϵ_{xc} functionals. (c) The variation in band gap of pristine supercell as a function of exact exchange fraction (α) contained in HSE06 ϵ_{xc} functional. 66
- 3.2 Ball and stick model of the optimized structures of (a) N_O , (b) N_i , (c) $(\text{N}_2)_\text{O}$, (d) Mn_Sr , (e) Mn_Ti , (f) Mn_i (g) $\text{Mn}_\text{Sr}\text{N}_\text{O}$, (h) $\text{Mn}_\text{Ti}\text{N}_\text{O}$, and (i) pristine SrTiO_3 69

3.3	2D projection of the 3D phase diagram that manifests the stable phases of (a) N-related, (b) Mn-related and (d) (N–Mn)-related charged defects having minimum formation energy as a function of μ_e and $\Delta\mu_O$. Here, on the x -axis, $\Delta\mu_O$ is varied according to T and p_{O_2} , and on the y -axis, μ_e is varied from the VBM to CBM of the pristine SrTiO ₃ . Colored regions show the most stable phases having a minimum formation energy at a given environmental condition. Top axes are showing the pressure (p_{O_2}) range at two temperatures: $T=300$ K and 1373 K. (c) Ball and stick model of the optimized structure of Mn _{Sr} N _O defect configuration.	70
3.4	Formation energy of N-related defects in SrTiO ₃ as a function of chemical potential of electron at (a) O-poor, (b) O-intermediate, and (c) O-rich conditions. Only those charge states of a particular defect are shown, which have lowest formation energies.	71
3.5	Formation energy of Mn-related defects in SrTiO ₃ as a function of chemical potential of electron at (a) O-poor, (b) O-intermediate, and (c) O-rich conditions.	72
3.6	Formation energy of (N–Mn)-related defects as a function of chemical potential of electron at (a) O-poor, (b) O-intermediate and (c) O-rich conditions.	73
3.7	Formation energy of (S–Mn)-related defect in SrTiO ₃ at (a) O-rich, (b) O-intermediate, and (c) O-poor conditions.	74
3.8	3D phase diagram that shows the most stable phases of (a) S–Mn, (b) S–Rh, and (c) N–Rh codoped SrTiO ₃ having minimum formation energy as a function of $\Delta\mu_O$ and μ_e	74
3.9	Electronic density of states for the supercell of (a) pristine SrTiO ₃ , (b) N _O , (c) Mn _{Sr} , (d) Mn _{Ti} , (e) Mn _{Sr} N _O , and (f) Mn _{Ti} N _O defect configurations.	76
3.10	Atom-projected partial density of states of (a) S _O , (b) Rh _{Ti} , (c) Rh _{Sr} , (d) Rh _{Ti} S _O , (e) Rh _{Sr} S _O , (f) Rh _{Ti} N _O , (g) Mn _{Sr} S _O , (h) Rh _{Sr} N _O , and (i) Mn _{Ti} S _O codoped SrTiO ₃	78
3.11	Spatially average (a) imaginary [Im (ϵ)] and (b) real [Re (ϵ)] part of the dielectric function for codoped SrTiO ₃ obtained using HSE06 ϵ_{xc} functional. Spatially average (c) imaginary [Im (ϵ)] and (d) real [Re (ϵ)] part of the dielectric function for monodoped SrTiO ₃	79

3.12	Spatially average (a) real (ϵ_1) and (b) imaginary (ϵ_2) part of the dielectric function obtained by $G_0W_0@HSE06$ for the pristine, (N/Mn) monodoped and (N–Mn) codoped $SrTiO_3$	80
3.13	Band-edge alignment of pristine, monodoped, and codoped $SrTiO_3$ w.r.t. water redox potential levels (H^+/H_2 , O_2/H_2O). The solid and dashed red lines in forbidden region represent the highest occupied and lowest unoccupied defect states, respectively. The highlighted ellipses indicate the most potent candidates for photocatalytic water splitting.	81
3.14	Band structure calculated using the HSE06 ϵ_{xc} functional of (a) pristine $SrTiO_3$, (b) $Mn_{Sr}N_O$, and (c) $Mn_{Ti}S_O$ codoped $SrTiO_3$	83
4.1	Atom-projected partial density of states (pDOS) of (a) pristine, (b) defective $CaTiO_3$ (O-vacancy at the O1 site, i.e., in the CaO plane), (c) defective $CaTiO_3$ (O-vacancy at the O2 site, i.e., in the TiO_2 plane), (d) crystal structure of orthorhombic (space group $Pbnm$) $CaTiO_3$	89
4.2	(a) Band-edge alignment of pristine and defective $CaTiO_3$ (bulk) and (b) the Gibbs free energy of formation (ΔG) for N_2 fixation over the (001) surface of pristine and defective $CaTiO_3$ (here, V_O represents single O-vacancy at the surface). Here, the second step of hydrogenation to form $N_2H_2^*$ is not considered over the pristine surface, as the first step of hydrogenation to form N_2H^* is endothermic.	90
5.1	(a) Structure of $Cs_2AgInCl_6$, and (b) Partial substitution with metals M(I), M(II), M(III) and with halogen X at Ag/In and Cl sites, respectively.	96
5.2	Radial distribution function of (a) $AgCl_6$ octahedral unit of $Cs_2AgInCl_6$, and (b) $CuCl_6$ octahedral unit of $Cs_2Cu_{0.25}Ag_{0.75}InCl_6$	97
5.3	Change in band gap on increasing the concentration of impurity atoms.	97
5.4	Decomposition energy (ΔH_D) for the decomposition of pristine and other configurations into binary compounds, and band gap using the ϵ_{xc} functionals (a) PBE and (b) HSE06. (c) Decomposition energy (ΔH_D) for decomposition into ternary compounds using HSE06 ϵ_{xc} functional.	100
5.5	Bandstructure of $Cs_2Au_{0.25}Ag_{0.75}InCl_6$ (a) without SOC, and (b) with SOC using HSE06 ϵ_{xc} functional.	105

5.6	Atom-projected pDOS using HSE06 ϵ_{xc} functional of (a) pristine $\text{Cs}_2\text{AgInCl}_6$, (b) $\text{Cs}_2\text{AgGa}_{0.25}\text{In}_{0.75}\text{Cl}_6$, (c) $\text{Cs}_2\text{Cu}_{0.25}\text{Ag}_{0.75}\text{InCl}_6$, (d) $\text{Cs}_2\text{Au}_{0.25}\text{Ag}_{0.75}\text{InCl}_6$, (e) $\text{Cs}_2\text{Zn}_{0.50}\text{Ag}_{0.75}\text{In}_{0.75}\text{Cl}_6$, (f) $\text{Cs}_2\text{Mn}_{0.50}\text{Ag}_{0.75}\text{In}_{0.75}\text{Cl}_6$, (g) $\text{Cs}_2\text{AgInBr}_{0.04}\text{Cl}_{5.96}$, and (h) $\text{Cs}_2\text{AgInI}_{0.04}\text{Cl}_{5.96}$	107
5.7	Atom-projected pDOS using HSE06 ϵ_{xc} functional of (a) $\text{Cs}_2\text{AgCo}_{0.25}\text{In}_{0.75}\text{Cl}_6$ and (b) $\text{Cs}_2\text{AgIr}_{0.25}\text{In}_{0.75}\text{Cl}_6$	108
5.8	Spatially average (a) imaginary [$\text{Im}(\epsilon)$] and (b) real [$\text{Re}(\epsilon)$] part of the dielectric function obtained by HSE06 for the pristine, and alloyed $\text{Cs}_2\text{AgInCl}_6$	108
5.9	Spatially average (a) imaginary [$\text{Im}(\epsilon)$] and (b) real [$\text{Re}(\epsilon)$] part of the dielectric function, (c) absorption coefficient and (d) band gap obtained by $G_0W_0@HSE06$ for the pristine $\text{Cs}_2\text{AgInCl}_6$ and other mixed sublattices.	109
5.10	Spatially average (a) imaginary [$\text{Im}(\epsilon)$] and (b) real [$\text{Re}(\epsilon)$] part of the dielectric function obtained by $G_0W_0@HSE06$ for the pristine, $\text{Cs}_2\text{AgCo}_{0.25}\text{In}_{0.75}\text{Cl}_6$, and $\text{Cs}_2\text{AgIr}_{0.25}\text{In}_{0.75}\text{Cl}_6$	110
5.11	Optical properties of (un)mixed $\text{Cs}_2\text{AgInCl}_6$: (a) refractive index (η), (b) extinction coefficient (κ), (c) reflectivity (R), (d) absorption coefficient (α), (e) optical conductivity (σ), and (f) energy loss spectrum (L) using $G_0W_0@HSE06$	112
6.1	Imaginary part of the electronic dielectric function with light polarization perpendicular to c -axis (ϵ_{xx}) for BaZrS_3 with different number of valence (NO) and conduction bands (NV) used in electron-hole interaction kernel.	117
6.2	(a) Model fitting for model-BSE (mBSE). (b) Spatially average imaginary [$\text{Im}(\epsilon)$] part of the dielectric function for BaZrS_3 with different k -mesh using mBSE. Imaginary part using GW-BSE is shown for reference by orange color. Calculated values of inverse of the static ion-clamped dielectric function $\epsilon_\infty^{-1} = 0.117$ and the screening length parameter $\lambda = 1.20$ are used in mBSE.	118
6.3	Imaginary part of electronic dielectric function for BaZrS_3 with light polarization along all three lattice vectors. For other chalcogenide perovskites as well, the minute anisotropy in dielectric function is existed.	118
6.4	Schematic crystal structure of orthorhombic (a) AZrS_3 ($A = \text{Ca}, \text{Sr}, \text{Ba}$) in distorted phase and (b) $\alpha\text{-SrZrS}_3$ in needle-like phase. Electronic pDOS of (c) CaZrS_3 , (d) $\beta\text{-SrZrS}_3$, (e) $\alpha\text{-SrZrS}_3$, and (f) BaZrS_3 using HSE06 ϵ_{xc} functional.	119

- 6.5 Electronic band structure of (a) CaZrS_3 , (b) $\alpha\text{-SrZrS}_3$, (c) $\beta\text{-SrZrS}_3$, and (d) BaZrS_3 using PBE ϵ_{xc} functional. 120
- 6.6 Imaginary [$\text{Im}(\epsilon)$] part of the dielectric function for BaZrS_3 with light polarization perpendicular to c-axis (ϵ_{xx}), obtained using different level of theories, specifically, PBE, HSE06, $G_0W_0@PBE$, $G_0W_0@HSE06$, $BSE@G_0W_0@PBE$, and $BSE@G_0W_0@HSE06$. First peak corresponds to the band gap of BaZrS_3 . . 122
- 6.7 Optical spectra of CaZrS_3 calculated using single-shot GW (G_0W_0) and self-consistent GW (scGW) on top of PBE orbitals. 123
- 6.8 Spatially averaged imaginary [$\text{Im}(\epsilon)$] part of the dielectric function for (a) CaZrS_3 , (b) $\alpha\text{-SrZrS}_3$, (c) $\beta\text{-SrZrS}_3$, and (d) BaZrS_3 obtained using $G_0W_0@PBE$ and $BSE@G_0W_0@PBE$. Peaks with turquoise color represent the oscillator strength. 124
- 6.9 Ionic contribution to the dielectric function for (a) CaZrS_3 , (b) $\alpha\text{-SrZrS}_3$, (c) $\beta\text{-SrZrS}_3$, and (d) BaZrS_3 obtained using DFPT. 126
- 6.10 Exciton binding energy (E_B) of (a) CaZrS_3 , (b) $\alpha\text{-SrZrS}_3$, (c) $\beta\text{-SrZrS}_3$, and (d) BaZrS_3 , as a function of dielectric constant. The intersection of the curve with vertical dashed line defines the upper bound obtained using the static electronic dielectric constant (at high frequency) and horizontal dashed line defines the lower bound obtained by the static ionic dielectric constant (at low frequency). 127
- 6.11 Spectroscopic limited maximum efficiency of AZrS_3 ($A = \text{Ca}, \text{Sr}, \text{and Ba}$). . . . 129

List of Tables

3.1	Band gap of different codopants in SrTiO ₃ using PBE ϵ_{xc} functional	63
3.2	The chemical potentials at different environmental conditions	68
3.3	Effective masses (in terms of free-electron mass m_e) at the band edge for pristine, Mn _{Sr} N _O , and Mn _{Ti} S _O codoped SrTiO ₃ . The masses m_{he} , m_{le} , m_{hh} , and m_{lh} correspond to heavy-electron, light-electron, heavy-hole, and light-hole bands, respectively.	82
5.1	Band gap evolution with respect to the number of bands using G ₀ W ₀ @PBE of Cs ₂ AgInCl ₆	95
5.2	Tolerance factor, octahedral factor, band gap, and decomposition energy (for decomposition into binary compounds) using PBE and HSE06 ϵ_{xc} functionals of different configurations	99
5.3	Decomposition energy (for the decomposition into ternary compounds) of Cs ₂ Cu _x Ag _{1-x} InCl ₆	104
5.4	Band gap (in eV) using PBE, HSE06, and G ₀ W ₀ @HSE06 for different configurations	106
5.5	The high frequency ‘ion-clamped’ dielectric constant (ϵ_∞) using G ₀ W ₀ @HSE06	111
6.1	Band gap (in eV) of CaZrS ₃ , α -SrZrS ₃ , β -SrZrS ₃ , and BaZrS ₃ using the PBE ϵ_{xc} functional	116
6.2	Band gap (in eV) of CaZrS ₃ , α -SrZrS ₃ , β -SrZrS ₃ , and BaZrS ₃ using G ₀ W ₀ @PBE with different number of bands	116
6.3	Calculated lattice parameters of AZrS ₃ (A = Ca, Sr, Ba) perovskites. The experimental values are provided in brackets. For distorted perovskites, specifically, CaZrS ₃ , β -SrZrS ₃ , and BaZrS ₃ , the experimental values are from Ref [2]. For α -SrZrS ₃ , the experimental values are from Ref [3]	120

6.4	Effective mass of electron, hole, and reduced mass (in terms of free-electron mass m_e) of chalcogenide perovskites along a $\Gamma-Z$ high-symmetry path	121
6.5	Band gap (in eV) of chalcogenide perovskites	121
6.6	Excitonic parameters for chalcogenide perovskites	125
6.7	Upper and lower bounds on exciton binding energy E_B for chalcogenide perovskites	125
6.8	Electron-phonon coupling parameters for chalcogenide perovskites	128
6.9	Polaron mobilities (μ) of CaZrS_3 , $\alpha\text{-SrZrS}_3$, $\beta\text{-SrZrS}_3$, and BaZrS_3 at $T = 300\text{ K}$	128